

0.015–0.6 mg/L NO<sub>2</sub>-N or 0.05–2.0 mg/L NO<sub>2</sub>

**APC 341**

**Scope and application:** For wastewater, drinking water, table water, surface water and mineral water.



## Test preparation

### Test storage

Storage temperature: 15–25 °C (59–77 °F)

### pH/Temperature

The pH of the water sample must be between pH 3–10.

The temperature of the water sample and reagents must be between 15–25 °C (59–77 °F).

### Before starting

Not more than **3 hours** should elapse between sampling and analyzing the sample.

Review safety information and expiration date on the package.

Review the Safety Data Sheets (MSDS/SDS) for the chemicals that are used. Use the recommended personal protective equipment.

Dispose of reacted solutions according to local, state and federal regulations. Refer to the Safety Data Sheets for disposal information for unused reagents. Refer to the environmental, health and safety staff for your facility and/or local regulatory agencies for further disposal information.

## Interferences

The ions listed in the table have been individually checked against the given concentrations and do not cause interference. The cumulative effects and the influence of other ions have not been determined. Chromium(VI) ions interfere with the determination. Copper(II) ions interfere with the determination even at concentrations below 1 mg/L.

The measurement results must be subjected to plausibility checks (dilute and/or spike the sample).

Interference level	Interfering substance
2000 mg/L	Cl <sup>-</sup> , SO <sub>4</sub> <sup>2-</sup>
1000 mg/L	K <sup>+</sup> , NO <sub>3</sub> <sup>-</sup>
500 mg/L	NH <sub>4</sub> <sup>+</sup> , PO <sub>4</sub> <sup>3-</sup> , Ca <sup>2+</sup>
100 mg/L	Mg <sup>2+</sup>
50 mg/L	Cr <sup>3+</sup>
25 mg/L	Co <sup>2+</sup> , Zn <sup>2+</sup> , Cd <sup>2+</sup> , Mn <sup>2+</sup> , Hg <sup>2+</sup>
12 mg/L	Ni <sup>2+</sup>
10 mg/L	Ag <sup>+</sup> , Fe <sup>2+</sup>
5 mg/L	Sn <sup>4+</sup> , Fe <sup>3+</sup>

## Summary of method

Nitrites react with primary aromatic amines in acidic solution to form diazonium salts. These combine with aromatic compounds that contain an amino group or a hydroxyl group to form intensively colored azo dyes.



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